

heterolytic bond-dissociation energy

The energy required to break a given bond of some specific compound by heterolysis. For the dissociation of a neutral molecule AB in the gas phase into A^+ and B^- the heterolytic bond-dissociation energy $D(A^+ B^-)$ is the sum of the bond dissociation energy, $D(A-B)$, and the adiabatic ionization energy of the radical $A\cdot$ minus the electron affinity of the radical $B\cdot$.

Source:

PAC, 1994, 66, 1077 (*Glossary of terms used in physical organic chemistry (IUPAC Recommendations 1994)*) on page 1121